The free energy of formation of carbon monoxide at 298.1°K. is −33,000 calories per mole as previously given. The entropy at 298.1°K. is increased by 0.003 to 47.316 calories per mole per degree."—J. O. CLAYTON AND W. F. GIAUQUE.

The Precision with which the Concentration of Solutions of Hydrochloric Acid and Sodium Hydroxide May Be Determined with the Immersion Refractometer. By E. Roger Washburn and Allen L. Olsen.

Page 3216. In the equation at the top of Table II the value "209" should be "229."—E. R. WASHBURN AND A. L. OLSEN.

Quantitative Determination of Calcium by the Magneto-optic Method. By Edna R. Bishop and C. B. Dollins.

Page 4586. In paragraph one, lines 8 and 14, the expression "$1/10^{n+2}$" should be "$1/10^{n}$."—E. R. BISHOP AND C. B. DOLLINS.

The Rate of Reaction between Chlorate and Sulfur Dioxide in Acid Solution. By A. C. Nixon and K. B. Krauskopf.

Page 4606. In the last line change "multiply" to "divide."

Page 4607. Omit Table I. In Table II change $\mu = 0.033$ to 0.33; change the fourth value of k from 2.57 to 2.84, and the fifth value from 2.57 to 2.60, giving a mean value of 2.72.

Page 4608. In the last line change 2.4 to 2.5.—A. C. NIXON AND KONRAD KRAUSKOPF.

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The Entropy of Polyatomic Molecules and the Symmetry Number. By Joseph E. Mayer, Stephen Brunauer and Maria Goeppert Mayer.

Pages 37 ff. The author writes "Dr. Louis S. Kassel has drawn my attention to an arithmetical error in the paper. In the calculations on acetylene, 4 degrees of freedom at 730 cm.$^{-1}$ contribute 1.1 E. U. at 298°K. The total absolute entropy should be 49.4 E. U. at 298°K. and the practical entropy 46.7, with a possible addition of the 1.1 E. U. due to the vibration at 730 cm.$^{-1}$."—JOSEPH E. MAYER.

Heat Capacity Curves of the Simpler Gases. I. Heat Capacity, Entropy and Free Energy of Gaseous Nitric Oxide from near Zero Absolute to 5000°K. By Herrick L. Johnston and Alan T. Chapman.

Pages 167 ff. "In Table IX, the 'free energy' value at 1125° should read 52.776. Other values in this table, as well as in Table VII of the second paper in this series [Johnston and Walker, THIS JOURNAL, **55**, 172 (1933)], deviate from the smooth curve of a large scale deviation plot by 0.001 or 0.002 units and in a few instances by as much as 0.004 unit but, with the correction given above, no values deviate by more than 0.004 unit. This is well below the limits of error (0.01 to 0.04 calories) imposed by other factors, including a small uncertainty in the value of the molar gas constant R."—H. L. JOHNSTON.